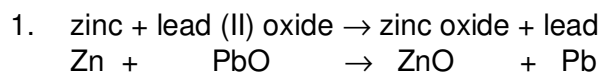
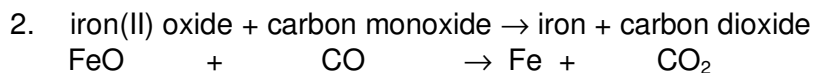


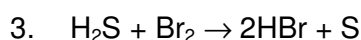
Name : _____ () Class: _____ Date : _____

Transfer of Oxygen

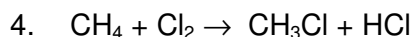
zinc is **oxidised** because it **gains oxygen**;
 lead(II) oxide is **reduced** because it **loses** oxygen.



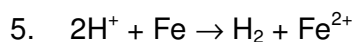
carbon monoxide is **oxidised** because it **gains** an extra oxygen to form carbon dioxide;
 iron(II) oxide is **reduced** because it **loses oxygen**.

Transfer of Hydrogen

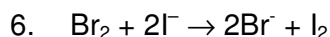
H_2S is **oxidised** because it **loses** hydrogen to become S;
 Br_2 is **reduced** because it **gains hydrogen** to become HBr.



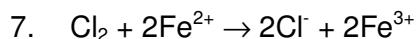
CH_4 is **oxidised** because it **loses** one hydrogen atom to become CH_3Cl ;
 Cl_2 is **reduced** because it **gains hydrogen** to become HCl.

Transfer of Electrons

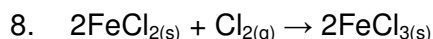
Fe is **oxidised** because it **loses** 2 electrons to form the Fe^{2+} ion;
 H^+ ions are **reduced** because they **gain 1 electron each** to form the hydrogen molecule, H_2 .



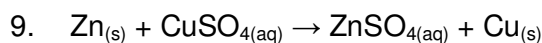
I^- ions are oxidised because they **lose** electrons to become the neutral I_2 molecule;
 Br_2 molecule is **reduced** because it **gains electrons** to form negative Br^- ions.



Fe^{2+} ions are **oxidised** because they **lose electrons** to become Fe^{3+} ions;
 Cl_2 molecule is **reduced** because it **gains** electrons to form negative Cl^- ions.
Note : each Fe^{2+} loses one electron

Change in oxidation state

Iron (II) chloride undergoes **oxidation** because oxidation state of iron **increases from +2 in FeCl_2 to +3 in FeCl_3** .
 Chlorine undergoes **reduction** as its oxidation state **decreases from 0 in Cl_2 to -1 in Cl^-** .



Zinc undergoes **oxidation** because its oxidation state **increases from 0 in Zn to +2 in ZnSO_4** .
 Copper (II) sulfate undergoes **reduction** as oxidation state of copper **decreases from +2 in CuSO_4 to 0 in Cu**.